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water by dissociation, dispersion, or ionization. 8.1.2 Describe how the physical properties of a solution can differ from those of its solute and solvent. 8.1.3 Identify energy changes that occur during the formation of a solution. 8.1.4 Describe factors affecting the

Section 8.1 Formation of Solutions (pages 228–234) This section explains the parts of a solution, the processes that occur when compounds dissolve, and how the properties of a solution compare with those of its solvent and solute. Reading Strategy (page 228) Comparing and Contrasting Contrast dissociation and ionization by listing

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Contrast dissociation and ionization by listing

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8.1.3 Identify energy changes that occur during the formation of a solution. 8.1.4

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**Solution:** The weight of potassium chloride required is 20% of the total weight of the solution, or  $0.2 \times 30\text{g} = 6.0\text{g}$  of KCl.

The remainder of the solution ( $30 - 6 = 24$ ) g consists of water.

Thus you would dissolve 6.0 g of KCl in 24 g of water.

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Volume of solution (in liters) Calculate the molar concentration (Molarity, M) of each of the following solutions: 16. 2.3 moles of sodium chloride in 0.45 liters of solution.  $M = \frac{2.3 \text{ moles}}{0.45 \text{ L}} = 5.1 \text{ M}$  17. 1.2 moles of calcium carbonate in 1.22 liters of solution.  $M = \frac{1.2 \text{ moles}}{1.22 \text{ L}} = 0.98 \text{ M}$  18. 0.09 moles of sodium sulfate in 12 mL of solution.

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