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Acid Base Calculations Answers

Solutions to Review
Problems for Acid/Base
Chemistry Answers: Acids

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& bases and pH
calculations: AS90310
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NOTES – CHAPTERS 20
AND 21 Acids and Bases ...
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(2004 ... Questions: Acids &
bases and pH calculations:

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Chemistry Made Easy
(PDF) Acid bases and salts
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CHEMISTRY NOTES –
CHAPTERS 20 AND 21

Acids and Bases ...

Solutions to Review

Problems for Acid/Base

Chemistry Answers: Acids
& bases and pH

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Acid/Base Calculations

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Practice Problems

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Chemistry Worksheet 11 a.

Hints for Working

Acid/Base Equilibrium ...

Acid Base Calculations

Answers -

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Calculations & Answers -
Science Skool!

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Problems for Acid/Base
Chemistry 8. Characterize
the solution formed
(strong acid, strong, base,
weak acid, weak base,
buffer or neutral) when
equal volumes of the
following are mixed.
Explain. $0.5 \text{ M NaOH} +$

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0.5 M $\text{HC}_2\text{H}_3\text{O}_2$ The
following reaction occurs:
 $\text{HC}_2\text{H}_3\text{O}_2 \dots$

Answers: Acids & bases
and pH calculations:
AS90310 (2004-2011 &
AS91166 (2012)
QUESTION (2012:3) (a)

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(i) Complete the table
below showing the
conjugate acids and bases.

Conjugate acid	Conjugate base
HCO_3^-	CO_3^{2-}
H_2O	OH^-
HCN	CN^-

(ii) $\text{HPO}_4^{2-}(\text{aq})$ Write
equations for the reactions
of HPO_4^{2-} with water:

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one where it acts as an
acid ...

Worksheet 21 - Acid/Base
Calculations 1. Start with
100.00 mL of a 0.10 M
solution of acetic acid, CH₃COOH. It has a pH =
2.87 at 25°C. Calculate

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the K_a of acetic acid at 25
°C. a) Write the acid
dissociation reaction. b)
Write an expression for K_a ,
in terms of ...

**ACID-BASE
CALCULATIONS (LIVE)
02 JUNE 2015 Section A:**

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Summary Notes Acids and

Bases - Calculations

Neutralisation Reaction

Neutralisation is the

reaction between an acid

and a base. The end point

is reached when the acid

and base are chemically

equivalent. An indicator is

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used during neutralisation
reactions to show the end
point.

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TASK 1 – Bronsted-

Lowry acids & bases 1

Acid = H_2O , base = NH_3

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2 Acid = HCl, base = H

2O 3 Acid = HCOOH,

base = KOH 4 Acid =

HCl, ... Buffer solution

calculations 1 a 3.35 b

4.72 c 4.77 2 a 4.35 b 3.57

3 a 47.9 g b 0.0113 g 4 a

pH ...

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separately; the best we can do is compare two different acid-base systems, and determine the extent to which the bases are able to compete against each other for the proton. 2. Water itself can act both as an acid and a base, and most

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of the practical
applications of acid-base
chemistry are those
involving aqueous
solutions.

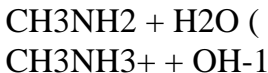
Acid and Base Worksheet -
Answers. 1) Using your
knowledge of the

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Brønsted-Lowry theory of acids and bases, write equations for the following acid-base reactions and indicate each conjugate acid-base pair: a) $\text{HNO}_3 + \text{OH}^- \rightarrow \text{H}_2\text{O} + \text{NO}_3^-$. HNO_3 and NO_3^- make one pair OH^- and H_2O

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make the other. b)



Calculate the pH of an unbuffered 0.010M acetic acid solution in which 0.004 M of OH^- is being added from an basic

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source. $\text{CH}_3\text{COOH} + \text{OH}^- \rightarrow$
 $\text{CH}_3\text{COO}^- + \text{H}_2\text{O}$

Acids and Bases -

Neutralization Goals : To
gain an understanding of :

1. The properties of acids
and bases
2. pH and pOH
calculations
3. Definitions

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of acids and bases 4.

Neutralization reactions

NOTES: Properties of
acids : • Form hydrogen
ions (H^+ or protons) in
solution • Form

hydronium (a water
molecule bonded to a
hydrogen ion shown as H

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Solutions to Review
Problems for Acid/Base
Chemistry 8. Characterize
the solution formed
(strong acid, strong, base,
weak acid, weak base,
buffer or neutral) when
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Explain. $0.5 \text{ M NaOH} +$
 $0.5 \text{ M HC}_2\text{H}_3\text{O}_2$ The
following reaction occurs:
 $\text{HC}_2\text{H}_3\text{O}_2 + \text{OH}^- \rightarrow \text{C}_2\text{H}_3\text{O}_2^- + \text{H}_2\text{O}$

Answers: Acids & bases
and pH calculations:
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AS91166 (2012)

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QUESTION (2012:3) (a)

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H_2O	OH^-
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(ii) Write equations for the reactions
 $\text{HPO}_4^{2-}(\text{aq})$

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of HPO_4^{2-} with water:
one where it acts as an
acid ...

(a) Identify three
conjugate acid-base pairs
in the equations above. (b)
 HCO_3^- can act as both
an acid and a base. Specify

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which equation above (1 or 2) shows HCO_3^- acting as an acid. Give a reason for your answer.

QUESTION (2004:3) Two acids of the same concentration, hydrochloric acid (HCl) and propanoic acid (CH_3

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CH₃COOH), have

Worksheet 21 - Acid/Base
Calculations 1. Start with
100.00 mL of a 0.10 M
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COOH. It has a pH =
2.87 at 25°C. Calculate
the K_a of acetic acid at 25

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oC. a) Write the acid
dissociation reaction. b)
Write an expression for K_a ,
in terms of ...

Worksheet: Acid base
problems - AP level
Problems 1 - 10 Problem
#1: Calculate the pH of the

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solution that results when
40.0 mL of 0.100 M NH_3
is: a) Diluted with 20.0
mL of distilled water. b)
Mixed with 20.0 mL of
0.200 M HCl solution.

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TASK 1 – Bronsted-

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4.72 c 4.77 2 a 4.35 b 3.57

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3 a 47.9 g b 0.0113 g 4 a
pH ...

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at California State
University, East Bay. Acid
Base Practice Worksheet
Calculations Answer each

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of the following to the best
of your ability and check
the

Acid and Base Worksheet -
Answers. 1) Using your
knowledge of the
Brønsted-Lowry theory of
acids and bases, write

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equations for the following
acid-base reactions and
indicate each conjugate
acid-base pair: a) $\text{HNO}_3 + \text{OH}^- \rightarrow \text{H}_2\text{O} + \text{NO}_3^-$.
 HNO_3 and NO_3^- make
one pair OH^- and H_2O
make the other. b)
 $\text{CH}_3\text{NH}_2 + \text{H}_2\text{O} \rightarrow$

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 $\text{CH}_3\text{NH}_3^+ + \text{OH}^-$

Solution 19: (a) A common acid base indicator and a universal indicator: Acid base indicator like litmus tells us only whether a given substance is an acid or a

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base. Universal indicator gives an idea as to how acidic or basic a substance is. An universal indicator gives different colours with solutions of ...

Acids and Bases -
Neutralization Goals : To

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gain an understanding of :

1. The properties of acids and bases
2. pH and pOH calculations
3. Definitions of acids and bases
- 4.

Neutralization reactions

NOTES: Properties of acids : • Form hydrogen ions (H^+ or protons) in

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solution • Form
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molecule bonded to a
hydrogen ion shown as H

Solutions to Review
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Chemistry 8. Characterize
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(strong acid, strong, base,
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Explain. $0.5 \text{ M NaOH} +$
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Answers: Acids & bases
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QUESTION (2012:3) (a)

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Conjugate acid Conjugate

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base HCO_3^- CO_3^{2-} H_2O OH^- HCN CN^- (ii)

HPO_4^{2-} (aq) Write

equations for the reactions

of HPO_4^{2-} with water:

one where it acts as an

acid ...

Worksheet 21 - Acid/Base

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Calculations 1. Start with 100.00 mL of a 0.10 M solution of acetic acid, CH_3COOH . It has a $\text{pH} = 2.87$ at 25°C . Calculate the K_a of acetic acid at 25°C . a) Write the acid dissociation reaction. b) Write an expression for K

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a, in terms of ...

p3 Recognizing Acid/Base
Properties when Ionics are
Dissolved in Water p11
pH Calculations;
Relationships between pH
and pOH p4 Answers p12
K_a: Sense + Calculations.

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Using K_a or pK_a to
Calculate $[H^+]$ and/or pH;
using pH to calculate K_a
or pK_a p5 Conceptual
Questions. Acids, Bases,
and Conjugates,
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Lowry acids & bases 1

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base = KOH 4 Acid =

HCl , ... Buffer solution

calculations 1 a 3.35 b

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4.72 c 4.77 2 a 4.35 b 3.57
3 a 47.9 g b 0.0113 g 4 a
pH ...

Equivalent strong base and
weak acid A-or B $[\text{OH}^-] =$
(K bCb) $1/2$ 0.1M NaOH
and 0.1M acetic acid
Examples that use the acid

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buffer equation Weak acid

and conjugate weak base

HA and A $[H^+] = K$

a_{Ca}/C_b 0.2 M acetic acid

and 0.1M sodium acetate

Strong acid and weak base

HA and A $[H^+] = K_a$

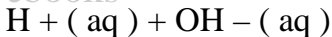
a_{Ca}/C_b 0.2 M HCl and 0.4

M sodium acetate

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Examples that use the ...

**Acid Base Calculations
Answers** Page 1 of 6 Acid-
Base Calculations The Ion-
Product Constant for
Water, K_w Water
undergoes ionization to a
small extent: $H_2O(l) ?$

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The equilibrium constant
for the reaction is the ion-
product constant for water
 $K_w \dots$

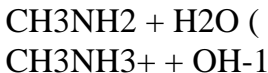
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make the other. b)



Solution 19: (a) A
common acid base
indicator and a universal
indicator: Acid base
indicator like litmus tells

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us only whether a given substance is an acid or a base. Universal indicator gives an idea as to how acidic or basic a substance is. An universal indicator gives different colours with solutions of ...

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Titration Calculations &
Answers 1. Use the
information to determine
the concentration of the
hydrochloric acid. • A 25
cm³ sample of
hydrochloric acid is
sucked into a pipette and
transferred into a 250 cm³

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volumetric flask. The solution is made up to the mark. • 25 cm³ of the diluted acid is transferred into a conical flask using a pipette.

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supplementary mindset.
Yeah, right of entry mind is
one that is needed gone
reading the book. You may
as well as craving to pick
what instruction and lesson
that is useful for you or
harmful. But in fact, this
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of many people to live, but

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as a consequence new features that will save you to come up with the money for perfection.

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