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# **Answer Key For Mixed Gas Law Problems**

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Mixed Gas Laws Worksheet - Solutions 1) How many moles of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K?  $n = \frac{PV}{RT} = \frac{(2.8 \text{ atm})(98 \text{ L})}{(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(292 \text{ K})} = 11$  moles of gas 2) If 5.0 moles of  $\text{O}_2$  and 3.0 moles of  $\text{N}_2$

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are placed in a 30.0 L tank at a temperature of 25 0

WS 5.5: Mixed Gas Law Problems. Directions: Solve the following problems. Round your answers using significant figures. 1)

Calculate the mass of 15.0 L of  $\text{NH}_3$  at  $27^\circ\text{C}$  and 900.0 mm Hg .

2) A volume of 26.5 mL of nitrogen gas was collected in a tube at a temperature of  $17^\circ\text{C}$  ...

Mixed Gas Law Problems A diver blows a 0.75 L bubble

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underneath the water. As it rises to the surface, the pressure goes from 2.25 atm to 1.03 atm. What will be the volume of the air in the bubble at the surface? The pressure in a car tire is 1.88 atm at 250C.

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Boyle's Law Problems: 1. If 22.5 L of nitrogen at 748 mm Hg are compressed to 725 mm Hg at constant temperature. What is the

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from 2.25 atm to 1.03 atm. What will be the volume of the air in the bubble at the surface? The pressure in a car tire is 1.88 atm at 250C.

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ANSWER KEY Boyle's, Charles' and Gay-Lussac's Gas Problems  
1. If a gas occupies 2.60 liters at a pressure of 1.00 atm, what will

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be its volume at a pressure of 3.50 atm? 0.743 L (Boyle's Law) 2. A gas occupies 900.0 mL at a temperature of 27.0 °C. What is the volume at 132.0 °C? 1215 mL (Charles' Law) 3.

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