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The chemical equation contains no phases, so we must use bond energies to solve this problem. This means we start with Lewis structures. Using those and Table 13.1, we know that 2 Cl-Cl bonds (each worth 240 kJ/mole) and 4 H-O bonds (each worth 459 kJ/mole) must be broken while 4 H-Cl bonds (each worth 428 kJ/mole) and 1 O=O bond (worth 494 kJ/mole) must be formed.

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