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**ELECTROLYTES** Remember a solution is a homogeneous mixture where a solute is dissolved in a solvent. Aqueous solutions are solutions where the solvent is water.

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Properties of Aqueous Solutions •  
electrolytes--solutions ...

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OF SOLUTIONS ... CO<sub>2</sub> in aqueous solution is 0.031 mol/L • atm at 25°C.  
before = 0.16 mol/L after =  $1.2 \times 10^{-5}$  mol/L ... solution and interferes with the

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IMFs of the solvent). There are fewer solvent molecules on the surface to escape.

**AP CHEMISTRY NOTES 10-4**

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**AQUEOUS EQUILIBRIA: SOLUBILITY  
PRODUCT** Saturated solutions of  
“insoluble” salts are another type of  
chemical equilibria. For a saturated solution

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of AgCl, the equation would be:  $\text{AgCl (s)} \rightleftharpoons \text{Ag}^+ + \text{Cl}^-$  The solubility product constant (K

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Applications of Aqueous Equilibria 0 2 4 6  
8 10 12 14 0 5 10 15 20 25 pH Volume of  
NaOH (aq) added (mL) pH vs mL of NaOH  
Added Trial 1 23) A 25.0-mL sample of a

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solution of an unknown compound is titrated with a 0.115 M NaOH solution...

The osmotic pressure of the solution is 7.55

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atm at 25°C. The molecular weight of nicotine is \_\_\_\_\_ g/mol. A)43 B)0.60 C)160 D)50 E)28 26) A solution is prepared by dissolving 6.00 g of an unknown

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nonelectrolyte in enough water to make  
1.00 L of solution. The osmotic pressure of  
this solution is ...

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**STOICHIOMETRY** Chemists have used reactions to produce the materials Teflon, nylon, Dacron, Kevlar, polystyrene, and PVC among many others! **WATER, THE**



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**COMMON SOLVENT • Properties –**  
Water has a high specific heat, high heat of vaporization, high cohesive and adhesive

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**AP Chemistry Unit #4 (Key) Chapter 4 –  
Zumdahl & Zumdahl Types of Chemical  
Reactions & Solution Stoichiometry**

**Students should be able to: Predict to some**

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extent whether a substance will be a strong electrolyte, weak electrolyte, or nonelectrolyte.

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Solution Stoichiometry Lecture Outline 4.1  
General Properties of Aqueous Solutions A  
solution is a homogeneous mixture of two**

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or more substances. A solution is made when one substance (the solute) is dissolved in another (the solvent). The solute is the



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substance that is present in smallest amount.

AP Chemistry Chapter 4. Aqueous  
Reactions and Solution Stoichiometry - 3 -

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Sample Exercise 4.4 (p. 132) Write the molecular, total ionic and net ionic equation for the precipitation reaction that occurs when solutions of calcium chloride and

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sodium carbonate are mixed. Practice  
Exercise 1 (4.4)

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a homogeneous mixture where a solute is  
dissolved in a solvent. Aqueous solutions

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are solutions where the solvent is water.

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aqueous solution of pure benzoic acid is titrated using standardized 0.150 M NaOH.  
(a) After addition of 15.0 mL of the 0.150 M NaOH, the pH of the resulting solution is

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4.37. Calculate each of the following. (i)  
[H<sup>+</sup>] in the solution [H<sup>+</sup>] = 10<sup>-4.37</sup> M =  
4.3 × 10<sup>-5</sup> M

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## AP Chemistry — CHAPTER 16 STUDY ...

•This definition is quite narrow in scope as it limits us to aqueous solutions. 16.2

Brønsted-Lowry Acids and Bases •We can

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use a broader, more general definition for acids and bases that is based on the fact that acid-base reactions involve proton transfers.

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## Zumdahl & Zumdahl Types of Chemical Reactions & Solution Stoichiometry

Students should be able to: Predict to some extent whether a substance will be a strong

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electrolyte, weak electrolyte, or nonelectrolyte.

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or more substances. A solution is made when one substance (the solute) is dissolved in another (the solvent). The solute is the

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substance that is present in smallest amount.

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Anderson - MCHS Solution Stoichiometry  
AP Chem. – Unit 02 Page 5 Chapter 4:  
Reactions in Aqueous Solutions Common  
Student Misconceptions Chapter 4.

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## Aqueous Reactions and Solution

Stoichiometry • Molarity is moles of solute per liter of solution, not per liter of solvent.

- Students sometimes use moles instead of

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molarity in M initial V

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Solutions 36. In aqueous solution, a chloride ion is attracted to which end of the water molecule? (A) the hydrogen end, which is the partial positive pole (B) the



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hydrogen end, which is the partial negative pole (C) the oxygen end, which is the partial positive pole

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Chemical Transformations ... A reactant  
that does not react B. Insoluble salt that  
falls out of solution C. The product that is**

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still aqueous (aq) D. None of the above  
Question 3 When solutions of two ionic compounds are combined and a ...

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