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Atomic Structure and Chemical Bonding
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Lewis structures (electron dot) show valence electrons of an atom as dots
Hydrogen has one dot, representing its 1s electron
Carbon has four dots ($2s^2 2p^2$) due to 4 e⁻ in valence shell
Kekulé structures (line-bond structures) have a line drawn

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between two atoms indicating a 2 e-covalent bond. Stable molecule results at completed shell, octet (eight dots) for main-group

This model proposed that atom is considered as a uniform positively charged sphere and electrons are embedded in it. An important feature of Thomson model of an atom was that mass of atom is considered to be evenly spread over the atom. Thomson model of atom is also called as Plum pudding, raisin pudding or watermelon model. Thomson model of atom was discarded because it could not explain ...

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