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**SECTION 4 Date CHAPTER 11 REVIEW Gases Class SHORT ANSWER Answer the following questions in the space provided.**

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Increasing rate of effusion. List the following gases in order (Assume all gases are at the same temperature and pressure.)  
(c) HCl (d) 2. Explain your reasoning for the order of gases you chose in item 1 above. Refer to the kinetic-

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Gas Calculations 11.3 Equation  
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between temperatures in the Celsius and  
Kelvin scales. (Section 2.6) Describe the  
particle nature of gases. (Section 3.1)  
Convert the amount of one substance in

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(water softening and purification), and in  
the separation and removal of gases and  
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agriculture, animal husbandry and  
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**SHORT ANSWER** Answer the following questions in the space provided. c c The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol.



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(b) divided by the mass of 1 mol.  $nRT$  (c) multiplied by 22.4 L. (d) divided by 22.4 L.  
For the expression  $V = (a)$

c. Area of contact =  $6.5 \text{ cm}^2$  Pressure = force / area = = Gases and Pressure In Chapter 10, you read about the kinetic-molecular theory of matter. You were also introduced to how this theory explains some of the properties of ideal gases. In this chapter, you will study the predictions of kinetic-molecular theory for gases in more detail.

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SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. The molar mass of a gas at STP

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is the density of that gas (a) multiplied by the mass of 1 mol. (c) multiplied by 22.4 L. (b) divided by the mass of 1 mol. (d) divided by 22.4 L. 2. For the expression  $V = \frac{nRT}{P}$

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partial = total each gas =  $RT \ Pn \ V$  To get a

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the mass of 1 mol. (c) multiplied by 22.4 L.  
(b) divided by the mass of 1 mol. (d)  
divided by 22.4 L. 2. For the expression  $V$   
 $nR P T$

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