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Stoichiometry CHAPTER 11

SOLUTIONS MANUAL Section 11.1

Defining Stoichiometry pages 368–372

Practice Problems pages 371–372 1.

Interpret the following balanced chemical equations in terms of particles, moles, and mass. Show that the law of conservation of mass is

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11.1.1 Objectives Describe the types of relationships indicated by a balanced chemical equation. State the mole ratios from a balanced chemical equation. Review Vocabulary reactant: the starting substance in a chemical reaction New Vocabulary stoichiometry mole ratio Defining

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Stoichiometry Stoichiometry CHAPTER 11 SOLUTIONS MANUAL Section 11.1 Defining Stoichiometry pages 368–372 Practice Problems pages 371–372 1.

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We can see from the stoichiometry of the reaction that $\frac{3}{2}$ mol of O_2 is required to produce 1 mol of H_2SO_4 . This is a standard stoichiometry problem of the type presented in Section 11.4, except this

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rancher.budee.org Q. $4\text{NH}_3 + 5\text{O}_2 \rightarrow 4\text{NO}$
 $+ 6\text{H}_2\text{O}$. What is the total number of

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moles of H_2O produced when 12 mole of NH_3 is Chapter 11.5: Stoichiometry Involving Gases - Chemistry ... ANSWER KEY 4 49.3% Rh 23.4 % C 27.3 % N 5. a) PBr_5 b) $\text{Zr}(\text{BO}_3)_2$ Zirconium (VI) borate 6.

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12. 160.10 g 13. 160.10 g 14. They are equal. 15. A mole ratio is a ratio between ...

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