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*Chemistry, 11e(Brown/LeMay/Bursten/Murphy) Chapter 15 Chemical Equilibrium
15.5 Algorithmic Questions 1) Phosphorous trichloride and phosphorous
pentachloride equilibrate in the presence of molecular chlorine according to the
reaction: $PCl_3(g) + Cl_2(g) \rightleftharpoons PCl_5(g)$ An equilibrium mixture at 450 K contains
 $PPCl_3 = 0.202 \text{ atm}$ $PCl_2 = 0.256 \text{ atm}$, and $s PPCl_5 = 3.45 \text{ atm}$.What is the value of
 K_p at this ...*

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Chapter 15 Chemical Equilibrium

(a) Write the equilibrium-constant expression for K_c for the following reaction: (b) With the information given in sample exercise 15.3, determine the value of this equilibrium constant at 25 °C. b. a. Writing products over reactants, we have
SAMPLE EXERCISE 15.3 Interpreting the Magnitude of an Equilibrium ...

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AP Chemistry Practice Test: Ch. 15 – Applications of Aqueous Equilibria 0 2 4 6 8 10 12 14 0 5 10 15 20 25 pH Volume of NaOH (aq) added (mL) pH vs mL of NaOH Added Trial 1 23) A 25.0-mL sample of a solution of an unknown compound is titrated with a 0.115 ...

19R - ch. 15 practice test - AP site - AP Chemistry 15 ...

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8/3/2017 · AP Chemistry 15 Chemical Equilibrium PRACTICE TEST 1.

Consider the reaction system, $\text{CoO}(s) + \text{H}_2(g) \rightleftharpoons \text{Co}(s) + \text{H}_2\text{O}(g)$. The equilibrium constant expression is a) $\frac{[\text{CoO}][\text{H}_2]}{[\text{Co}][\text{H}_2\text{O}]}$ d) $\frac{[\text{H}_2]}{[\text{H}_2\text{O}]}$ b) $\frac{[\text{Co}][\text{H}_2\text{O}]}{[\text{CoO}][\text{H}_2]}$ e) $\frac{[\text{H}_2\text{O}]}{[\text{H}_2]}$ c) $\frac{[\text{Co}][\text{H}_2\text{O}]}{[\text{H}_2]^2}$.

Chemistry Chapter 15 - Chemical Equilibrium Flashcards ...

expresses, for any reaction, the relationship between the concentrations of the reactants and products present at equilibrium. suppose we have the general equilibrium: $aA + bB \rightleftharpoons dD + eE$ where A, B, D, E are the chemical species involved and a, b, d, and e are their coefficients in the balanced chemical equations. according to the law of mass action, the equilibrium ...

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Chapter 15 – Chemical Equilibrium 15.1 The Concept of Equilibrium Figure: 13.1 from Chemistry by McMurray & Fey ? Figure 13.1(a) 24() 2()g 2 g colorless brown NO NO ?? we start with reactant, N₂O₄, so the solution is colorless

Chapter 15 Lecture- Chemical Equilibrium - SlideShare

25/1/2010 · Chapter 15 Lecture- Chemical Equilibrium. 1. Chapter 15 - Chemical Equilibrium Sections 15.1 - 15.4 Equilibrium & the Equilibrium Constant (K) 2. The Concept of Equilibrium Chemical equilibrium occurs when a reaction and its reverse reaction proceed at the same rate. 3.

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Testname: CH_13_PRAC_TEST_EQUILIBRIUM.TST MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question. 1) A ID: chem9b 15.1-1 2) C ID: chem9b 15.1-7 3) A ID: chem9b 15.1-12 4) A ID: chem9b 15.1-15 5) B ID: chem9b 15.1-18 6) E ID: chem9b 15.1-24 7) A ID: chem9b 15.1-27 8) B ID: chem9b 15.1-31 9) B ID: chem9b 15.1-35 10) E

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Q. Use the K_a values to determine whether there are more products or reactants at equilibrium for the given reaction. $K_a \text{ NH}_4^+ = 5.6 \times 10^{-10}$ $K_a \text{ HCO}_3^- = 4.8 \times 10^{-11}$

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AP Chemistry Practice Test: Ch. 15 – Applications of Aqueous Equilibria

0	2	4	6	8	10	12	14	0	5	10	15	20	25
pH													
Volume of NaOH (aq) added (mL)													
pH vs mL of NaOH Added													
Trial 1													

23) A 25.0-mL sample of a solution of an unknown compound is titrated with a 0.115 ...

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Chapter 15 Chemical Equilibrium

*Chapter 15 Chemical Equilibrium * Note: On the AP exam, the required question has always been on equilibrium. All possible types of equilibrium will be discussed in chapters 15,16,17. Throughout these chapters, I will be giving you past AP equilibrium questions. Chemical equilibrium: The condition in a reaction when the concentrations of reactants*

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Consider the reaction system $\text{CoO}(s) + \text{H}_2(g) \rightleftharpoons \text{Co}(s) + \text{H}_2\text{O}(g)$ 4 a 4.45×10^{-4} b 6.67*

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Chapter 15 Practice Questions

15) At equilibrium, _____. A) all chemical reactions have ceased B) the rate constants of the forward and reverse reactions are equal C) the rates of the forward and reverse reactions are equal D) the value of the equilibrium constant is 1 E) the limiting reagent has been consumed 15) 4

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Chapter 15 Practice Questions

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a) Catalysts do not alter the position of equilibrium: they do not shift the equilibrium to the left or right b) At equilibrium, $G = 0$ c) If G for a reaction is negative, the forward reaction happens spontaneously

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12. At equilibrium a) All chemical reactions have ceased. b) $K_c = \frac{\text{rate}_{\text{reverse}}}{\text{rate}_{\text{forward}}}$ c) The rate constant for the forward reaction equals the rate constant of the ... Microsoft Word - Chem 1220 Recitation Activity Chapter 15 Practice Exam Questions.docx ...

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Assume that the only effect of a catalyst on the reaction is to lower the activation energies of the forward and reverse reactions, as shown in Figure 15.15. Using the Arrhenius equation (Section 14.5), prove that the equilibrium constant is the same for the catalyzed reaction as for the uncatalyzed one.

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