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### Chapter 15 Review Acids Bases Answer Key

454 CHAPTER 15 Formula Acid name HF hydrofluoric acid HCl hydrochloric acid HBr hydrobromic acid HI hydriodic acid H<sub>2</sub>S hydrosulfuric acid TABLE 15-1 Names of Binary Acids FIGURE 15-2 A strip of pH paper dipped into vinegar turns red, showing that vinegar is an acid.

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of 6 CHAPTER 15. ACIDS AND BASES  
General Properties of Acids & Bases: Acids  
Bases taste sour taste bitter act corrosive  
feel slippery turn litmus paper red turn  
litmus paper blue examples: vinegar, lemon  
juice, gastric juice, soft drinks ...

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What we will learn: • Bronsted acids and  
bases • Acid-base properties of water • pH -  
a measure of acidity • Strength of acids and  
bases • Weak acids (bases) and acid (base)  
ionization constant • Diprotic and  
polyprotic acids • Molecular structure and  
strength of acids • Acid-base properties ...

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stomach produce hydrochloric acid, HCl  
(aq) to kill unwanted bacteria to ...

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**CHAPTER 15 REVIEW Acids and Bases**  
**SECTION 15-1 SHORT ANSWER** Answer the following questions in the space provided. 1. Name the following compounds as acids: a.  $\text{H}_2\text{SO}_4$  b.  $\text{H}_2\text{SO}_3$  c.  $\text{H}_2\text{S}$  d.  $\text{HClO}_4$  e. hydrogen cyanide 2. Which (if any) of the acids mentioned in item 1 are

**CHAPTER 15 REVIEW Acid-Base Titration and pH** **SECTION 1 SHORT ANSWER** Answer the following questions in the space provided. 1. Calculate the following values without using a calculator.

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1.810 M a. The  $[H_3O^+]$  is  $1 \times 10^{-6}$  M in a solution. Calculate the  $[OH^-]$ .  
1.510 M b. The  $[H_3O^+]$  is  $1 \times 10^{-9}$  M in a solution. Calculate the  $[OH^-]$ .  
1.102 M c.

### GCh15-1 Chapter 15. Acids and Bases

What we will learn:

- Bronsted acids and bases
- Acid-base properties of water
- pH - a measure of acidity
- Strength of acids and bases
- Weak acids (bases) and acid (base) ionization constant
- Diprotic and polyprotic acids
- Molecular structure and strength of acids
- Acid-base properties ...

Chapter 15 Acids and Bases Notes page 1 of 6  
**CHAPTER 15. ACIDS AND BASES**  
General Properties of Acids & Bases: Acids  
Bases taste sour taste bitter act corrosive



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feel slippery turn litmus paper red turn  
litmus paper blue examples: vinegar, lemon  
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Calendar CHAPTER 15 REVIEW Acids  
and Bases SECTION 15-1 SHORT

ANSWER Answer the following questions  
in the space provided 1 Name the following  
compounds as acids: a  $\text{H}_2\text{SO}_4$  b  $\text{H}_2\text{SO}_3$   
c  $\text{H}_2\text{S}$  d  $\text{HClO}_4$  e hydrogen cyanide 2  
Which (if any) of the acids mentioned in  
item 1 are Guide to Chapter 15.

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### CHAPTER 15: ACIDS AND BASES

Problems: 1-10, 15-18, 51-54, 63-64, 67-68, 71-72

#### 15.1 PROPERTIES OF ACIDS &

**BASES** Acids Bases – produce hydrogen

ions,  $H^+$  – produce hydroxide ions,  $OH^-$  –

taste sour – taste bitter; feel soapy, slippery

– turn blue litmus paper red – turn red

litmus paper blue

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### CHAPTER 15 REVIEW Acid-Base

#### Titration and pH SECTION 1 SHORT

ANSWER Answer the following questions

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in the space provided. 1. Calculate the following values without using a calculator.

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stomach produce hydrochloric acid, HCl  
(aq) to kill unwanted bacteria to ...

### CHAPTER 15: ACIDS AND BASES

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is a strong acid,  $[H_3O^+] = 7.0 \times 10^{-6} =$   
 $5.15 \times 10^{-6} = [H_3O^+] = 7.0 \times 10^{-6} M.$

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Behind Lucy, The Chimp Raised To

4 CHAPTER 15–16 REVIEW:: ACIDS  
AND BASES KEY 15.139 True or False?  
If false explain why the statement is wrong  
(a) All Lewis acids are Brønsted acids:  
False. A Lewis acid such as CO<sub>2</sub> is not a  
Brønsted acid. It does not have a hydrogen

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ion to donate. (b) The conjugate base of an acid always carries a negative charge. False.

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that in water, but correlations and contrasts  
between solvents are also presented.  
Fundamental background material is  
provided in the initial chapters: quantitative



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