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questions in the space provided. 1. _____ In metals, the valence electrons are considered to be (a) attached to particular positive ions. (c) immobile. (b) shared by all surrounding atoms.(d) involved in covalent bonds. 2.

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more epoxide groups per monomer, and this molecule contains a tight C—O—C ring structure.

Chapter 6: Review. Bonding Comparison Chart. IONIC
COVALENT METALLIC Types of Atoms Involved (Metal,
Nonmetal) Metals and nonmetals Nonmetals Metals and metals
Method of Bond Formation (Valence Electrons) ... Come Together:
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bond actually looks like! Chapter 9 - Covalent Bonding I. Covalent Bonding: attractive force produced as a result of shared (pgs 189-193) electrons. A. A _____ is formed when two or more atoms bond covalently. B. Bonds form when there is a balance of attractive and repulsive forces between two atoms: C. Single Covalent Bond & Lewis Structures

6.1 Ionic Bonding (5 Questions and Answers in Complete has a positive or negative charge and acts as a unit is a 6 Chemical Bonds

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Chemical Bonding tabulate the geometry of some molecules
showing sp, sp², sp³, dsp², and dsp³ hybridisation; explain the
formation of and bonds in CH₄, C₂H₄ and C₂H₂; explain
molecular orbital theory; write the molecular orbital configuration

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of H₂, N₂, O₂ and F₂ molecules; define bond length and bond order and relate them and

Modern Chemistry 47 Chemical Bonding CHAPTER 6 REVIEW
Chemical Bonding SECTION 4 SHORT ANSWER Answer the following questions in the space provided. 1. _____ In metals, the valence electrons are considered to be (a) attached to particular positive ions. (c) immobile. (b) shared by all surrounding atoms.(d) involved in covalent bonds. 2.

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Nonmetal) Metals and nonmetals Nonmetals Metals and metals
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189electrons. -193) A. A _____ is formed when two or more atoms
bond covalently. B. Bonds form when there is a balance of
attractive and repulsive forces between two atoms: C. Single

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When H^+ forms a bond with H_2O to form the hydronium ion H_3O^+ , this bond is called a coordinate covalent bond because _____. a. both bonding electrons come from the oxygen atom b. it forms an especially strong bond c. the electrons are equally shared d. the oxygen no longer has eight valence electrons ____ 28.

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formation of and bonds in CH_4 , C_2H_4 and C_2H_2 ; explain molecular orbital theory; write the molecular orbital configuration of H_2 , N_2 , O_2 and F_2 molecules; define bond length and bond order and relate them and

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