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6-1 CHAPTER 6 THERMOCHEMISTRY: ENERGY FLOW AND CHEMICAL CHANGE 6.1 The sign of the energy transfer is defined from the perspective of the system. Entering the system is positive, and leaving the system is negative. 6.2 No, an increase in temperature means that heat has been transferred to the surroundings, which makes  $q$  positive.

6-1 CHAPTER 6 THERMOCHEMISTRY: ENERGY FLOW AND CHEMICAL CHANGE CHEMICAL CONNECTIONS BOXED READING PROBLEMS B6.1 Plan: Convert the given mass in kg to g, divide by the molar mass to obtain moles, and convert moles to kJ of energy. Sodium sulfate decahydrate will transfer 354 kJ/mol. Solution: Heat (kJ) =  $3 \times 24 \times 2 \times 24 \times 2 \times 24 \times 2 \times 4 \times 2$

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1 Chem 1036 Chapter 6 Chapter 6 Thermochemistry: Energy Flow and Chemical Change Thermodynamics: - the study of energy changes in a system. -the study of the transformation of energy from one form to another. Thermochemistry:-the branch of thermodynamics that focuses on the heat involved in chemical (reactions) and physical changes.

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Energy Change Thermodynamic system: An isolated system, i.e. the system of interest such as molecules being surrounded by surrounding, and the surrounding is enclosed by an insulated wall so that no heat, no work, no molecules escape from the surrounding. The system can exchange heat and/or work to the surrounding. The energy of the system, internal energy which consists of molecular ...

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Thermochemistry: Energy Flow and Chemical Change 6.1 Forms of Energy and Their Interconversion 6.2 Enthalpy: Heats of Reaction

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c) The chemical reaction is absorbing energy. d) The chemical reaction is exothermic. e) More than one of these. ANS: c) The chemical reaction is absorbing energy. PAGE: 6.2 27. What is the heat capacity of mercury if it requires 167 J to change the temperature of 15.0 g mercury from 25.0°C to 33.0°C? a)  $6.92 \times 10^{-3} \text{ J/g}^\circ\text{C}$  b)  $1.12 \times 10 \dots$

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increase the potential energy of the system. In an endothermic process, the products have higher potential energy (weaker bonds on average) than the reactants. 3. First law of thermodynamics: the energy of universe is constant. A system can change its internal energy by flow of work, heat, or both ( $E = q + w$ ). Whenever a property is added to ?

Energy Change Thermodynamic system: An isolated system, i.e. the system of interest such as molecules being surrounded by surrounding, and the surrounding is enclosed by an insulated wall so that no heat, no work, no molecules escape from the surrounding. The system can exchange heat and/or work to the surrounding.. The energy of the system, internal energy which consists of molecular ...

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6.1 ?  $E = q + w$  The sign of the energy transfer is defined from the perspective of the system. 6.2 No. An increase in temperature

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means that heat has been transferred to the surroundings, which makes  $q$  positive. 6.3  $\Delta E = q + w = w$ , since  $q = 0$ .

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