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Step 1, Write down your given equation.

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For this example, you will use:  $C_3H_8 + O_2$   
 $\rightarrow H_2O + CO_2$  This reaction occurs when  
propane ( $C_3H_8$ ) is burned in the presence  
of oxygen to produce water and carbon



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**Chemistry: Chemical Word Equations KEY**  
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equation for each of the word equations below. 1. aqueous sodium chloride reacts with aqueous lead (II) nitrate to yield a lead (II) chloride precipitate and aqueous sodium

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nitrate

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?^ A^? ?#K " II c the final examination will be weighted as follows modules 1 –3 15 –20% modules 4 –6 80 –85% the format of the examination will be as follows: Part a:

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Chemical Equations Questions 1. Early Chinese alchemists discovered an early form of gunpowder. What was the composition of this substance? 2. What later

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developments were made to the gunpowder recipe that improved its pyrotechnic properties? 3.

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**Chemistry: Chemical Word Equations KEY**  
**Directions: Write a balanced chemical equation for each of the word equations below.**

1. aqueous sodium chloride reacts

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with aqueous lead (II) nitrate to yield a lead (II) chloride precipitate and aqueous sodium nitrate

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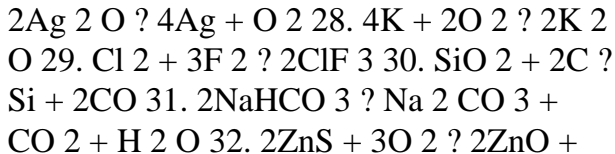
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**2SO 2 33. SO 2 ...**

**Chemistry: Quantitative Relationships in  
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chemical equation, we are satisfying the law of conservation of mass; that is, we are making sure that there are the same number of atoms of each element on both sides of

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the equation.

**Balancing Equations: Practice Problems 1.**  
Balance each of the following equations. (a)

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Fe + Cl<sub>2</sub> ??? FeCl<sub>3</sub> (b) Fe + O<sub>2</sub> ??? Fe<sub>2</sub>O<sub>3</sub>  
(c) FeBr<sub>3</sub> + H<sub>2</sub>SO<sub>4</sub> ??? Fe<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> + HBr  
(d) C<sub>4</sub>H<sub>6</sub>O<sub>3</sub> + H<sub>2</sub>O ??? C<sub>2</sub>H<sub>4</sub>O<sub>2</sub> (e) C<sub>2</sub>H<sub>4</sub>  
+ O<sub>2</sub> ??? CO<sub>2</sub> + H<sub>2</sub>O (f) C<sub>4</sub>H<sub>10</sub>O + O<sub>2</sub> ???

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$\text{CO}_2 + \text{H}_2\text{O (g)}$   $\text{C}_7\text{H}_{16} + \text{O}_2$  ???  $\text{CO}_2 +$   
 $\text{H}_2\text{O (h)}$   $\text{H}_2\text{SiCl}_2 + \text{H}_2\text{O}$  ???  $\text{H}_8\text{Si}_4\text{O}_4 +$   
 $\text{HCl (i)}$   $\text{HSiCl}_3 + \text{H}_2\text{O}$  ?? ...

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chemical equations for any five of eight  
given sets of chemical reactions. The new  
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developments were made to the gunpowder recipe that improved its pyrotechnic properties? 3.

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**Chemistry: Chemical Word Equations KEY**  
**Directions: Write a balanced chemical equation for each of the word equations below.**

1. aqueous sodium chloride reacts

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with aqueous lead (II) nitrate to yield a lead (II) chloride precipitate and aqueous sodium nitrate

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continued 2 21 linn Scientific nc ll ihts  
esere Sweet 16 Chemistry of Gases  
Tournament Supplementary Chemical

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Equations Worksheet Predict the products and balance the following chemical equations to identify the contestants in the Chemistry of Gases tournament. 1.Na 2



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$\text{CO}_3^{2-}(\text{s}) + \text{HCl}(\text{aq})?$

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- ANSWERS 26.  $2\text{Mg} + \text{Cl}_2 \rightarrow 2\text{MgCl}_2$  27.  
 $2\text{Ag}_2\text{O} \rightarrow 4\text{Ag} + \text{O}_2$  28.  $4\text{K} + 2\text{O}_2 \rightarrow 2\text{K}_2\text{O}$   
29.  $\text{Cl}_2 + 3\text{F}_2 \rightarrow 2\text{ClF}_3$  30.  $\text{SiO}_2 + 2\text{C} \rightarrow$   
 $\text{Si} + 2\text{CO}$  31.  $2\text{NaHCO}_3 \rightarrow \text{Na}_2\text{CO}_3 +$

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$\text{CO}_2 + \text{H}_2\text{O}$  32.  $2\text{ZnS} + 3\text{O}_2 \rightarrow 2\text{ZnO} + 2\text{SO}_2$  33.  $\text{SO}_2 \dots$

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**Chemical Equations** When we balance a chemical equation, we are satisfying the law of conservation of mass; that is, we are making sure that there are the same number

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of atoms of each element on both sides of the equation.

**Balancing Equations: Practice Problems 1.**

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Balance each of the following equations. (a)  
 $\text{Fe} + \text{Cl}_2 \rightarrow \text{FeCl}_3$  (b)  $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$   
(c)  $\text{FeBr}_3 + \text{H}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + \text{HBr}$   
(d)  $\text{C}_4\text{H}_6\text{O}_3 + \text{H}_2\text{O} \rightarrow \text{C}_2\text{H}_4\text{O}_2$  (e)  $\text{C}_2\text{H}_4$



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+ O<sub>2</sub> ??? CO<sub>2</sub> + H<sub>2</sub>O (f) C<sub>4</sub>H<sub>10</sub>O + O<sub>2</sub> ???  
CO<sub>2</sub> + H<sub>2</sub>O (g) C<sub>7</sub>H<sub>16</sub> + O<sub>2</sub> ??? CO<sub>2</sub> +  
H<sub>2</sub>O (h) H<sub>2</sub>SiCl<sub>2</sub> + H<sub>2</sub>O ??? H<sub>8</sub>Si<sub>4</sub>O<sub>4</sub> +

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HCl (i)  $\text{HSiCl}_3 + \text{H}_2\text{O} ?? \dots$

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