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Point Elevation . Go to the following website and perform the virtual experiment: IA State Chem Department- Colligative Properties. Part 2. Vapor Pressure.

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Experiment 1: Colligative  
Properties Determination  
of the Molar Mass of a  
Compound by Freezing  
Point Depression.

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**Objective:** The objective of this experiment is to determine the molar mass of an unknown solute by measuring the freezing point depression of a solution of this solute in a solvent as compared to the freezing point of the pure

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**Solutions Background:**

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properties of a solvent,  
such as freezing point

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**Experiment 1: Colligative  
Properties Determination  
of the Molar Mass of a  
Compound by Freezing  
Point Depression.**

**Objective: The objective  
of this experiment is to**

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determine the molar mass of an unknown solute by measuring the freezing point depression of a solution of this solute in a solvent as compared to the freezing point of the pure solvent.



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**Solutions Background:**

Colligative properties are  
properties of a solvent,

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depression and boiling  
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Colligative Properties

Prior Knowledge

Questions (Do these

BEFORE using the

Gizmo.) 1. When the

forecast calls for freezing

rain, salt is often spread on

the roads. Why is this

done? \_\_\_\_ It helps melt the

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ice on the roads and drive  
ways. \_\_\_\_ 2. Antifreeze  
is added to car radiators to  
prevent the water in the  
radiators from both  
freezing and boiling.

Pre-Lab Discussion .  
Distinguish between solute

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and solvent. Solute is the lesser component and solvent the greater component in a solution.

List three colligative properties and suggest a rationale for the choice of the word colligative to describe these properties.

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The three colligative properties are boiling point, freezing point, and vapor ...

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Properties - Definition,

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Types, Examples ... The colligative properties include vapor pressure lowering, boiling point elevation, freezing point depression, and osmotic pressure. The vapor pressure ...

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Colligative Properties

Prior Knowledge

Questions (Do these

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Gizmo.) 1. When the forecast calls for freezing rain, salt is often spread on the roads. Why is this done? \_\_\_ It helps melt the ice on the roads and drive ways. \_\_\_\_\_ 2. Antifreeze is added to car radiators to prevent the water in the

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radiators from both  
freezing and boiling.

These colligative  
properties include vapor  
pressure lowering, boiling  
point elevation, freezing  
point depression, and  
osmotic pressure. This



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small set of properties is of  
central importance to  
many natural phenomena  
and technological  
applications, ... Answers to  
Chemistry End of Chapter  
Exercises. 2.

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– Freezing Point

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Introduction. The physical properties of solutions that depend on the number of dissolved solute particles and not their specific type are known as colligative

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properties. These include  
freezing point depression,  
osmotic pressure, and  
boiling point elevation.

Freezing Point Depression  
and Colligative Properties  
Materials 1/2 cup milk 1/2  
cup whipping cream

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(heavy cream) or  
(substitute 4 Tbsp. half-n-  
half) 1/4 cup sugar 1/4  
teaspoon vanilla or vanilla  
flavoring (vanillin) 1/2 to  
3/4 cup sodium chloride  
(NaCl) as table salt or rock  
salt 2 cups ice 1-quart  
Ziploc™ bag 1-gallon

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1. Design experiments to answer a research question about the influence adding a solute has to the solvent's physical properties: freezing point and boiling point. 2. What

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influence does adding  
more solute to a solvent  
have on the freezing point  
and boiling point ...

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