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Ionic Bonds 9.1 Describing Ionic Bonds An ionic bond is a chemical bond formed by the electrostatic attraction between positive and negative ions. The bond forms between two atoms when one or more electrons are transferred from the valence shell of one atom to the valence shell of the other.

Chapter 9 Theories of Chemical Bonding 9-3 9-3 A covalent bond is the result of the overlap of orbitals on adjacent atoms. The bonding region is the location between the atomic nuclei, where electrons occupy the overlapping orbitals. For example, consider the covalent bond in hydrogen, H_2 (Figure 9.2). H: $1s$ Figure 9.2
The covalent bond in H_2

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Covalent bonding diagram for meth- ane (CHO. Note that both the

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carbon and hydrogen atoms have full outer shells. A bond can now form between the negatively-charged Cl^- and the positively-charged Na^+ . This type of bond is called an ionic bond. Ionic bonds typically form between one metal and one non-metal ion. The above reaction can be written as:

Chapter 9 Theories of Chemical Bonding 9-3 9-3 A covalent bond is the result of the overlap of orbitals on adjacent atoms. The bonding region is the location between the atomic nuclei, where electrons occupy the overlapping orbitals. For example, consider the covalent bond in hydrogen, H_2 (Figure 9.2). H: 1s Figure 9.2 The covalent bond in H_2

two atoms constitutes a covalent bond. The Lewis Model of Chemical Bonding. Covalent Bonding in H_2 H. H Two hydrogen atoms, each with 1 electron, can share those electrons in a covalent bond. $\text{H}:\text{H}$ •Sharing the electron pair gives each hydrogen an electron configuration analogous to helium.

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Virtual Ionic vs Covalent Bond Properties Lab ... As you watch the lab on youtube, fill out the data chart below. Answer the questions that follows. Compound Name Compound Formula Melting Point Dissolves in Water ... job is to identify whether the unknown

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compound is held together with ionic or covalent bonds using the evidence below.

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