

Enthalpy And Entropy Of A Borate Solution

ENTHALPY AND ENTROPY OF A BORAX SOLUTION (PDF)

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Enthalpy and Entropy of a Borax Solution Revised 4/28/15 $3 \ln K_{sp} = \ln (4(1.35)^3) = 2.29$ When graphing $\ln K_{sp}$ he would use inverse Kelvin temperature, $3.01 \times 10^{-3} \text{ K}^{-1}$. In today's experiment,

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students will prepare 5 saturated borax solutions at temperatures between

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Temperature of borate mixture in: (t1) _____ oC and T 1 _____ K Remember the volume of the borate solution titrated is 10.00 mL. Do not include the 20 mL of distilled water added in the titration. ΔH_0 of Borax Dissolution Evaluate the enthalpy change (ΔH_0) for borax dissolution using equation 5 where K 2 is

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The enthalpy change (ΔH°) will be determined from a plot of $\ln K_{sp}$ versus $1/T$, and the entropy change (ΔS°) will be calculated from equation (2). PRE-LAB QUESTIONS (1) Define: saturated solution, enthalpy, entropy, free energy, and spontaneity. (2) In your textbook, read the section on "Free Energy" and complete the table below.

Temperature of borate mixture in: (t1) _____ oC and T 1 _____ K
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 ΔH° of Borax Dissolution Evaluate the enthalpy change (ΔH°) for borax dissolution using equation 5 where K_2 is

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Tutorial 2 - Enthalpy and Entropy - Solutions Page 5 8. A process in which entropy increases and enthalpy increases will (go to completion/ reach a state of equilibrium/not occur at all) ___reach a state of equilibrium___ 9.

Temperature of borate mixture in: (t1) _____ oC and T 1 _____ K
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state of equilibrium____ 9.

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