

Flow Of Energy Heat And Work Answers

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- 3 - INTERNAL ENERGY, HEAT AND WORK c) radiation is the flow of heat from one place to another by means of electromagnetic wave – see Y6 Questions: 8. Why does paper wrapping keep hot things hot and cold things cold? 9. Some people say that fur coats would keep their owners warmer when worn inside out. Do you agree? 10.

The energy that flows from a warm object to a cool object 1. is called . The energy stored within the structural units of 2. chemical substances is called chemical . The study of heat 3. changes during chemical reactions and changes of state is called 4.. Heat energy can be measured in . This unit is 5.

energy and work –1st law of thermodynamics When heat flows into (or out of) a system, the system gains (or loses) an amount of energy equal to the amount of heat transferred. ?Heat = ?Internal

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Energy + Work – Adding heat to a system can: increase the internal energy of the system enable the system to do external work (or both)

Energy, Work, and Heat • Energy is the capacity to do work or to supply heat. Energy can be exchanged between objects by some combination of either heat or work: Energy = heat + work $E = q + w$ • Work is done when a force is exerted through a distance: Work = force \times distance • Heat is the flow of energy caused by a temperature difference.

4. A steady flow system loses 2 kW of heat also loses 4 kW of work. What is the net change in the energy of the fluid flowing through it? (-6 kW) 5. A steady flow system loses 3 kW of heat also loses 20 kW of work. The fluid flows through the system at a steady rate of 70 kg/s. The velocity at inlet is 20 m/s and at outlet it is 10 m/s. The ...

2. Heat is not the same as temperature, yet they are related. Explain how they differ from each other. Heat is a combination of kinetic energy (measured by temperature) and potential energy. a. Perform calculations using: ($q = m c \Delta T$) b. Determine if it's endothermic or exothermic 1. Gold has a specific heat of 0.129 J/(g \times °C). How

The SI unit of heat and energy is the , which is equal to 7. 0.2390 cal. The of a substance is the amount of heat it 8. takes to change the temperature of 1 g of the substance 1°C. Substances like , with low heat capacities, take a shorter time to heat up than substances with high heat capacities, such as .8 7 6 5 3 4 2 1 heat (joules or ...

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The study of heat 3. changes during chemical reactions and changes of state is called 4.. Heat energy can be measured in . This unit is 5. defined as the amount of heat needed to raise 1 g of water 1 °C. 6. The SI unit of heat and energy is the , which is equal to 7. 0.2390 cal. The of a substance is the amount of heat ...

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inside out. Do you agree? 10.

Chapter 2: Internal Energy, Work, Heat and Enthalpy 15 More
general formula for PV work, P does not need to be constant $\int P dV$
 $w = - \int P dV$ Sign Convention : Work done on the system raises
internal energy of system ($w > 0$) Work done by the system lowers
the internal energy ($w < 0$) Other forms of work: - electrical work $w = -q$
I Q is charge in coulombs I

Energy, Work, and Heat • Energy is the capacity to do work or to
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Heat is a combination of kinetic energy (measured by temperature)
and potential energy. a. Perform calculations using: ($q = m c \Delta T$) b.
Determine if it's endothermic or exothermic 1. Gold has a specific
heat of $0.129 \text{ J/(g}\cdot\text{°C)}$. How many joules of heat energy are
required to raise the temperature of 15 grams of gold from 22 °C to
 85 °C ?

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11. _____ A transfer of heat energy through space by means of electromagnetic waves. 12. _____ The study of the flow of heat energy. 13. _____ This states that heat will always flow from hot objects to cold objects, efficiency is always less than 100%, and total disorder can never decrease. 14.

displacement of a fluid, will be referred to as flow work. If an amount of fluid of volume V is inserted into a system against a pressure P , the work required to accomplish this is PV . Enthalpy, therefore, can be viewed as the sum of the internal energy of this fluid volume added to the system plus the flow work ...

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15/8/2020 · Energy is measured in terms of its ability to perform work or to transfer heat. Mechanical work is done when a force f displaces an object by a distance d : (1) $w = f \times d$. The basic unit of energy is the joule. One joule is the amount of work done when a force of 1 newton acts over a distance of 1 m; thus $1 \text{ J} = 1 \text{ N}\cdot\text{m}$.

17/5/2015 · The Internal Energy (ΔU or ΔE) Work, w , is the mechanical transfer of energy from one thing to another Heat, q , is the energy transferred from a hot object to a cold one upon contact.

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All energy transfers can be classified as either heat or work. . 11.
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(in) ?w (out) ?q ...

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