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basic. However, I may be somewhat surprised to discover that aqueous solutions of some salts (for example, sodium nitrate,  $\text{NaNO}_2$ , and potassium acetate,  $\text{KC}_2\text{H}_3\text{O}_2$ ) are basic, whereas others (for ...

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a neutralization reaction of acids and bases, and are almost always  
strong electrolytes. A hydrolysis reaction is the reaction that occurs  
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The salt hydrolysis formula of Sodium acetate is as follows:  
 $\text{CH}_3\text{COONa (aq)} \rightleftharpoons \text{CH}_3\text{COO}^- \text{ (aq)} + \text{Na}^+ \text{ (aq)}$  The acetate ion,  
thus formed, interacts with the hydrogen ion in water to form acetic  
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ion, thus formed, interacts with the hydrogen ion in water to form acetic acid and hydroxyl ion.

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