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Isotope Practice Worksheet Name:  
Learning Target: Use isotope notation to  
determine: element name/symbol, atomic

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number, number of electrons, number of neutrons, number of protons, mass number, atomic number, atomic mass. Isotope Notation: 1. Here are three isotopes ...

Answer: 126.86 amu 10. The natural abundance for boron isotopes is 19.9%  $^{10}\text{B}$  and 80.1%  $^{11}\text{B}$ . Calculate boron's atomic mass. Answer: 126.86 amu 11. Hydrogen is 99%  $^1\text{H}$ , 0.8%  $^2\text{H}$ , and 0.2%  $^3\text{H}$ . Calculate its average atomic mass. Answer: 1.21 amu 12. Rubidium is a soft, silvery-white metal that has two common isotopes,  $^{85}\text{Rb}$  and  $^{87}\text{Rb}$ . If the ...

**PART I.** Answer the questions based on the above reading. 1. What is an isotope?  
Isotopes are versions of the same element.

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They have the same number of protons and electrons as the element but different mass numbers and number of neutrons. 2. What does the number next to isotopes signify? The number indicates the isotope's mass number. 3.

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Atoms and Isotopes Worksheet Fill in the table with the correct information Isotope Isotope notation Atomic # Protons Electrons neutrons Oxygen-16 8 8 8 8 Bromine-80 35 35 35 45 Uranium-235 92 92 92 143 Copper-64 29 29 29 35 1. Describe the general arrangement of subatomic particles in the atom

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Use isotopic notation to determine: element name/symbol, atomic number, number of



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electrons, number of neutrons, number of protons, mass number, atomic number, and atomic mass. Here are three isotopes of an element:  ${}^6_{12}\text{C}$   ${}^6_{13}\text{C}$   ${}^6_{14}\text{C}$ . The element is: \_\_\_\_\_ The number 6 refers to the \_\_\_\_\_

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three isotopes of an element:  ${}^6_{12}\text{C}$   ${}^6_{13}\text{C}$   ${}^6_{14}\text{C}$  a. The element is: Carbon b. The number 6 refers to the atomic number c. The numbers 12 ... Balancing Chemical Equations – Answer Key Balance the equations below: 1)  $1 \text{ N}_2 + 3 \text{ H}_2$

22/10/2014 · Model 1 that supports your answer. YES, CARBON-12 AND CARBON-13 ARE ISOTOPES OF CARBON AND HAVE THE SAME ATOMIC NUMBER. b. Do all isotopes of an element have the same mass number? Give at least 1 example or counter-example from Model 1 that supports your answer. NO, CARBON-12 AND CARBON-13 ARE ISOTOPES OF CARBON BUT HAVE DIFFERENT MASS NUMBERS.

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12.

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Isotope notation Atomic # Protons  
Electrons neutrons Oxygen-16 8 8 8 8  
Bromine-80 35 35 35 45 Uranium-235 92

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92 92 143 Copper-64 29 29 29 35 1.

Describe the general arrangement of subatomic particles in the atom

Chemistry Worksheet NAME: \_\_\_\_\_

Isotope Notation Block ... 1. Uranium-235 and uranium-238 are considered isotopes of one another. How are uranium-235 similar, and how are they different? 2. Define isotope: 3. The number of protons in an atom is ... The isotope notation for nitrogen-15 is as follows: a. The number 15 is the ...

Use isotopic notation to determine: element name/symbol, atomic number, number of electrons, number of neutrons, number of protons, mass number, atomic number, and atomic mass. Here are three isotopes of an

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element:  ${}^6_{12}\text{C}$   ${}^6_{13}\text{C}$   ${}^6_{14}\text{C}$ . The element is:  
\_\_\_\_\_ The number 6 refers to the \_\_\_\_\_

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