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each statement or best answers each question. ... CH c. 3. VThat other information do you need in order to determine the molecu- ... Quiz Modern Chemistry 47 . Name Section Quiz, continued Class Date 6. Wthat is the empirical formula of a compound that contains 79.86% I

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29/3/2016 · Modern Chemistry 68 Chapter Test Name Class Date
Chapter Test A, continued _____ 21. Metal X replaces the ions of metal Y from solution, but it cannot replace the ions of metal Z from solution. The order these metals should have in the activity series (from ...

CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter
SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Explain the difference between the mass number and the atomic number of a nuclide. Mass number is the total number ...

3) 2, to form aluminum nitrate, $\text{Al}(\text{NO}_3)_3$? If so, write the balanced chemical equation for the reaction. Yes; because aluminum is above copper in the activity series, aluminum metal will replace copper in copper(II) nitrate. $2\text{Al}(s) + 3\text{Cu}(\text{NO}_3)_2(aq) \rightarrow 2\text{Al}(\text{NO}_3)_3(aq) + 3\text{Cu}(s)$
70 CHEMICAL EQUATIONS AND REACTIONS
MODERN CHEMISTRY

Modern Chemistry 99 Chapter Test Chapter: Gases PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. _____ 1. If the temperature of a gas remains constant, then the pressure of the gas will increase if the a. mass of the gas molecules decreases.

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6/4/2016 · Modern Chemistry 110 Chapter Test Name Class Date
Chapter Test B, continued _____ 7. Effervescence is the a.
dissolving of a gas in a liquid. b. escape of a liquid from a liquid-
liquid solution. c. escape of a solid from a solid-liquid solution. d.
escape of a gas from a gas-liquid solution. _____ 8.

Modern Chemistry 99 Chapter Test Chapter: Gases PART I In the
space provided, write the letter of the term or phrase that best
completes each statement or best answers each question. _____ 1. If
the temperature of a gas remains constant, then the pressure of the
gas will increase if the a. mass of the gas molecules decreases.

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below to answer the questions in this Chapter Test. In the space
provided, write the letter of the term or phrase that best completes
each statement or best answers ...

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CHAPTER 10 REVIEW States of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Match description on the right to the correct crystal type on the left. b ionic crystal (a) has mobile electrons in the crystal c covalent molecular crystal (b) is hard, brittle, and nonconducting a metallic crystal (c) typically has the lowest melting point of the four

Modern Chemistry 137 Chapter Test Name Class Date Chapter Test A, continued ____19. In the figure on the previous page, the pH at the equivalence point a. is equal to 7.0. b. is greater than 7.0. c. is less than 7.0. d. cannot be determined from the data given. ____20. In the figure on the previous page, the volume of titration standard

SECTION 3 Date CHAPTER 11 REVIEW Gases Class SHORT ANSWER Answer the following questions in the space provided. c The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (b) divided by the mass of 1 mol. nRT (c) multiplied by 22.4 L. (d) divided by 22.4 L. For the expression $V = (a)$ increasing P

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Modern Chemistry 137 Chapter Test Name Class Date Chapter Test A, continued _____19. In the figure on the previous page, the pH at the equivalence point a. is equal to 7.0. b. is greater than 7.0. c. is less than 7.0. d. cannot be determined from the data given. _____20. In the figure on the previous page, the volume of titration standard

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