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## Properties of Solutions 2 $\frac{3}{4}$

miscible—When two or more liquids mix  
(ex. Water and food coloring)  $\frac{3}{4}$

immiscible—When two or more liquids

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**DON'T** mix.—they usually layer if allowed to set for a while. (ex. Water and oil)

Concentration Units  $\frac{3}{4}$  Molarity (M) = # of moles of solute per liter of solution; IS

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temperature dependent.

Colligative Properties of Electrolyte  
Solutions (See A Closer Look on p. 494)

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For colligative properties, — Electrolyte solutions vary from those of nonelectrolyte solutions — electrolytes dissociate into ions in solutions ? increasing the total number of

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particles in solution

Properties of Solutions: Electrolytes and  
Non-Electrolytes c. Before testing the next

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solution, clean the electrodes by rinsing them liberally with distilled water from a wash bottle. Blot the outside of the probe end dry using a tissue. It is not necessary to

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dry the inside of the hole near the probe  
end. 6. Obtain the four Group B solution ...

Properties of Solutions: Electrolytes and



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Non-Electrolytes c. Before testing the next solution, clean the electrodes by rinsing them liberally with distilled water from a wash bottle. Blot the outside of the probe

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particles in solution

Properties of Solutions Chemistry, The  
Central Science , 10th edition Theodore L.

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**Brown; H. Eugene LeMay, Jr.; and Bruce E. Bursten. Solutions Solutions • Solutions are homogeneous mixtures of two or more pure substances. • In a solution, the solute is**



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dispersed uniformly throughout the solvent .

Reaction in a Bag Buffers and pH titrations  
(Chemistry Laboratory Previews)

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**Properties Of Solutions Lab Answers**  
Answers. Colligative properties are  
characteristics that a solution has that  
depend on the number, not the identity, of

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solute particles. In solutions, the vapor pressure is lower, the boiling point ...

## 4.1 General Properties of Aqueous

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**Solutions** • **Solution** - a homogeneous mixture – **Solute**: the component that is dissolved ? – **Solvent**: the component that does the dissolving Generally, the

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component present in the greatest quantity is considered to be the solvent. Aqueous solutions are those in ...

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## 8.3 Acid/Base Properties of Salt Solutions

Salts are solids (at SATP) composed of cations and anions arranged in a crystalline lattice. When dissolved in water, they

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dissociate into individual hydrated ions that may or may not affect the pH of the solution.

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19/3/2021 · 13: Properties of Solutions. In  
all solutions, whether gaseous, liquid, or

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solid, the substance present in the greatest amount is the solvent, and the substance or substances present in lesser amounts are the solute (s). The solute does not have to be in

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the same physical state as the solvent, but  
the physical state of the solvent usually ...

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(ex. Water and food coloring)  $\frac{3}{4}$   
immiscible—When two or more liquids  
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Concentration Units  $\frac{3}{4}$  Molarity (M) = # of moles of solute per liter of solution; IS temperature dependent.



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Central Science , 10th edition Theodore L.  
Brown; H. Eugene LeMay, Jr.; and Bruce  
E. Bursten. Solutions Solutions • Solutions**

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are homogeneous mixtures of two or more pure substances. • In a solution, the solute is dispersed uniformly throughout the solvent .

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**WORKSHEET:SOLUTIONS AND  
COLLIGATIVE PROPERTIES SET A: 1.**  
Find the molarity of all ions in a solution  
that contains 0.165 moles of aluminum

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chloride in 820. ml solution. Answer:  $[Al^{3+}] = 0.201\text{ M}$  ,  $[Cl^-] = 0.603\text{M}$ . 2. Find the molarity of each ion present after mixing 27 ml of 0.25 M  $HNO_3$  with 36 ml of 0.42 M

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Properties of some particular solutions 2 .  
Annex 1. Salt water solutions We study

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here basically aqueous solutions of common salt ( $\text{NaCl}$ ,  $=0.023+0.0355=0.0585$  kg/mol), i.e. M water / sodium-chloride liquid mixtures, called brines. Although the main

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motivation is the study of sea

Reaction in a Bag Buffers and pH titrations  
(Chemistry Laboratory Previews)

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characteristics that a solution has that  
depend on the number, not the identity, of



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solute particles. In solutions, the vapor pressure is lower, the boiling point ...

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to the properties of elements, bonding, solutions, as well as the basic principles of circuits.

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all solutions, whether gaseous, liquid, or

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solid, the substance present in the greatest amount is the solvent, and the substance or substances present in lesser amounts are the solute (s). The solute does not have to be in

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the same physical state as the solvent, but  
the physical state of the solvent usually ...

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