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Here are the heat capacities of the four substances: $0.10 \text{ cal/g } ^\circ\text{C}$, $0.25 \text{ cal/g } ^\circ\text{C}$, $1.0 \text{ cal/g } ^\circ\text{C}$, & $0.2 \text{ cal/g } ^\circ\text{C}$. Match & then label each substance with its specific heat capacity on the graph. See graph above. 7. If something has a high specific heat capacity will it take a lot of heat or a little heat to change its temperature? Explain ...

Worksheet- Calculations involving Specific Heat 1. For $q = m c \Delta T$: identify each

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variables by name & the units associated with it. q = amount of heat (J) m = mass (grams) c = specific heat ($\text{J/g}^\circ\text{C}$) ΔT = change in temperature ($^\circ\text{C}$)

2. Heat is not the same as temperature, yet ...

Here are the heat capacities of the four substances: $4.18 \text{ J/g } ^\circ\text{C}$, $1.00 \text{ J/g } ^\circ\text{C}$, $0.80 \text{ J/g } ^\circ\text{C}$, & $0.60 \text{ J/g } ^\circ\text{C}$. Match then each substance with its specific heat capacity on the graph. If something has a high specific heat capacity will it take a lot of heat or a little heat to change its

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choices Specific heat capacity is the amount of heat per unit mass required to raise the temperature of a substance by one Kelvin (or degree Celcius) Specific heat capacity is the same per unit mass for any substance.

7. Aqueous silver ion reacts with aqueous chloride ion to yield a white precipitate of solid silver chloride. When 10.0 mL of 1.00M AgNO_3 solution is added to 10.0mL of 1.00 M NaCl solution at 25oC in a calorimeter a white precipitate of AgCl forms and the temperature of the aqueous mixture increases to 32.6oC. Assuming that the specific heat of the aqueous mixture is 4.18 J/goC, that the ...

Latent heat and Specific heat capacity

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questions. 1. How much water at 50°C is needed to just melt 2.2 kg of ice at 0°C ? 2. How much water at 32°C is needed to just melt 1.5 kg of ice at -10°C ? 3. How much steam at 100° is needed to just melt 5 kg of ice at -15°C ? 4. A copper cup holds some cold water at 4°C .

Specific Heat Problems 1) How much heat must be absorbed by 375 grams of water to raise its temperature by 25°C ? 2) What mass of water can be heated from 25.0°C to 50.0°C by the addition of 2825 J? 3) What is the final temperature when 625 grams of water at 75.0°C loses 7.96×10^4 J?

Heat: ordinary level maths questions 2002

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Question 12 (b) [Ordinary Level] (i) Define specific heat capacity. (ii) -An electric kettle contains 1.5 kg of water. The specific heat capacity of water is $4180 \text{ J kg}^{-1} \text{ K}^{-1}$. Calculate the amount of energy required to raise the temperature of the water from 15°C to 100°C .

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key calculating ... Instead of writing the
answers to the questions on the ...

View Answer. The specific heat capacity of copper is $0.385 \text{ J g}^{-1} \text{ }^\circ\text{C}^{-1}$ whereas it is $0.128 \text{ J g}^{-1} \text{ }^\circ\text{C}^{-1}$

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for gold. Assume you place 100 g of each metal, originally at 25 ...

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Specific Heat Capacity specific heat capacity
tl fi nc au296r?!j 7t2 tet pc6f kl ti
(xt, how much heat is up 36 kg of hydrogen gas from 12.0 to

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