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Equilibrium and Reaction
Rates Study Guide Period:
Know the reaction
diagram 1. Label the
diagrams below as either
endothermic or exothermic
2. Locate the activation
energy 3. Which graph has

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its reactants at a higher
energy level than the
products? I 80 60 PE 20
QC Progress of the
reaction exothermic C-
Reaction Coordinate

C 23. reversible reactions
OL 24. chemical

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equilibrium Column A 25.

Le Chatelier's principle ,
Match each description in
Column B to the correct
term in Column A.

Column B a. state of
balance in which forward
and reverse reactions take
place at the same rate. b.

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When stress is applied to a system at equilibrium, the system changes to relieve the stress.

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27 As you can see, this
"molecular formula is not
very informative, it tells us
little or nothing about their
structure, and suggests that

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all proteins are similar,
which is confusing since ...

Chapter 14 Equilibrium

Notes page 4 of 6 14.3

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EQUILIBRIUM For the
reaction $\text{N}_2\text{O}_4(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$ Rate of forward

reaction = Rate of reverse
reaction $k_f[\text{N}_2\text{O}_4] = k_r$

$[\text{NO}_2]^2$ Rearrange: $[\text{N}_2\text{O}_4] = \frac{k_r}{k_f} [\text{NO}_2]^2$

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Choice Identify the choice

that best completes the

statement or answers

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reaction = Rate of reverse
reaction $k_f[\text{N}_2\text{O}_4] = k_r[\text{NO}_2]^2$ Rearrange: $[\] [$

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Equation 2 ALA] At
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describes a 2.

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reaction is in chemical
equilibrium when the rate
of its forward reaction
equals the rate of its
reverse reaction and the
concentrations of its
products and reactants
remain unchanged. The
chemical equation for the

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reaction at equilibrium is written using double arrows to indicate the overall reversibility of the reaction.

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what we're going to do is look at how the rate of the forward reaction changes if we put some pure N_2O_4 in a flask: If we put some pure N_2O_4 in a flask (No NO_2 yet!), there will be a

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high concentration of N_2O_4 . That is, there will be lots of N_2O_4 molecules to collide with each other.

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Note: Please ensure that the learners understand and know the factors affecting the rate of a reaction very well before attempting this section on chemical equilibrium. The only factors affecting chemical equilibrium are

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temperature, pressure and
concentration. 1.

Introduction: 5 minutes . 2.

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3 kinetics study of the rate
or speed at which
reactions occur a reaction
is the breaking and

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reforming of bonds to
make entirely new
compounds as products
reaction mechanism step
by step process needed to
make a product how you
get from a to b like a
recipe reactants

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Question 1. Equation 1
describes a 2.

Closed system, rate of
forward reaction equals
rate of reverse reaction,
constant temperature 12.
Describe how the forward
and reverse reaction rates

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are related once a reaction reaches equilibrium. Draw a graph to describe this relationship when a system is at equilibrium. Rate of forward reaction = rate of reverse reaction 13.

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reaction is in chemical equilibrium when the rate of its forward reaction equals the rate of its reverse reaction and the concentrations of its products and reactants remain unchanged. The chemical equation for the

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reaction at equilibrium is

written using double

arrows to indicate the

overall reversibility of the

reaction.

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what we're going to do is look at how the rate of the forward reaction changes if we put some pure N_2O_4 in a flask: If we put some pure N_2O_4 in a flask (No NO_2 yet!), there will be a high concentration of N_2O_4 . That is, there will

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be lots of N_2O_4 molecules
to collide with each other.

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Appendix 3.5B: Factors

Affecting the Rate of

Reactions: Lab Activity

(Answer Key) 26

Appendix 3.6A: Factors

Affecting the Rate of a

Reaction: Lab Activity 28

Appendix 3.6B: Factors

Affecting the Rate of a

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Reaction: Lab Activity
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4.1: Preparation of
Equilibrium Systems
(Demonstration) 3

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