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Chapter 6 Oxidation-Reduction Reactions

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UNIT 6 REDOX REACTIONS - A-Level Chemistry

Answers:!! Assigning!oxidation!numbers!!

1.

The!oxidation!number!of!hydrogen!is!+1!!Since!the!total!charge!is!0!in!the!compound,!and!no!net!electrical!charge!is!shown,!the!oxidation!number!of!sulfur!must!be!-2:!! +1!!J2!!!H

2S!!! 2.

The!oxidation!number!of!oxygen!is!-2.!With!no!net!charge!

By referring to your periodic table, predict the oxidation number of these elements when they form ions: a. -1 b. +4 c. +4 d. +1 e. +3 f. -2 g. +2 h. +1 i. +4 j. -1 k. +2 3.

Examine each of the following reactions and decide which are `redox' reactions.

Explain your choice: a. Cu changes to Cu^{2+}

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& Cl changes to Cl⁻. No changes in oxidation numbers – not a redox reaction.

Worksheet 25 - Oxidation/Reduction

Reactions Oxidation number rules:

Elements have an oxidation number of 0

Group I and II – In addition to the elemental oxidation state of 0, Group I has an oxidation state of +1 and Group II has an oxidation state of +2. Hydrogen –usually +1, except when bonded to Group I or Group II, when it forms hydrides, -1.

3.) The oxidation number of a monatomic ion is its ionic charge. In these cases the apparent charge of the atom is the real charge of the ion. a. For Group I elements (the alkali metals), the oxidation number is

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+1 in all compounds. Li^+ Na^+ K^+ Rb^+ Cs^+
+1 +1 +1 +1 +1 b. Likewise, for Group II elements (the alkaline earth metals), the oxidation number is

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Oxidation Number Oxidation numbers are a useful tool for determining whether a

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substance has been oxidized or reduced. An element that undergoes a change in oxidation number in the course of a reaction has been oxidized or reduced. Let's learn how to assign oxidation numbers.

28/6/2010 · Chapter 20 Worksheet: Redox ANSWERS I. Determine what is oxidized and what is reduced in each reaction.

Identify the oxidizing agent and the reducing agent, also. 1. $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$ Sr 0 to Sr^{2+} ; oxidized/reducing agent O^0 to O^{2-} ; reduced/ox. ag. 2. $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$ Li 0 to Li^{1+} ; oxidized/red. ag. S^0 to S^{2-} ; reduced/ox. ag. 3.

8/5/2013 · Redox practice worksheet Name: _____ Date: _____ 1. In which substance is the oxidation

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number of nitrogen zero? A. NH_3 B. N_2 C. NO_2 D. N_2O 2. What is the oxidation number of carbon in NaHCO_3 ? A. +6 B. +2 C. 4 D. +4 3. In the reaction $\text{Al}^0 + \text{Cr}^{3+} \rightarrow \text{Al}^{3+} + \text{Cr}^0$, the reducing agent is A. Al^0 B. Cr^{3+} C. Al^{3+} D. Cr^0 4. In the reaction $2\text{K} + \text{Cl}_2 \rightarrow 2\text{KCl}$, the species oxidized is A. Cl_2 B. Cl C. K D. K^+ 5.

Step 1 Assign an oxidation number to each atom in the equation and determine whether it is a redox reaction. The oxidation numbers of sulfur and oxygen change during the reaction. It is a redox reaction. Step 2 Identify the atom or atoms that undergo an increase in oxidation number and the atom or atoms that undergo a

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decrease in oxidation number.

By referring to your periodic table, predict the oxidation number of these elements when they form ions: a. -1 b. +4 c. +4 d. +1 e. +3 f. -2 g. +2 h. +1 i. +4 j. -1 k. +2 3.

Examine each of the following reactions and decide which are 'redox' reactions.

Explain your choice: a. Cu changes to Cu^{2+} & Cl changes to Cl^- b. No changes in oxidation ...

Oxidation Number Oxidation numbers are a useful tool for determining whether a substance has been oxidized or reduced. An element that undergoes a change in oxidation number in the course of a reaction has been oxidized or reduced. Let's learn

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how to assign oxidation numbers.

Ch 10 Oxidation and reduction 4(19)
elements is the number of charges possessed by that atom. The oxidation numbers of Fe, Fe²⁺ and Fe³⁺ are 0, +2 and +3, respectively In order to extend the concept of oxidation number to polyatomic molecules, it is necessary to

8/5/2013 · B. reduction, only C. both oxidation and reduction D. neither oxidation nor reduction 23. In the reaction $\text{AgNO}_3(\text{aq}) + \text{NaCl}(\text{aq}) \rightarrow \text{NaNO}_3(\text{aq}) + \text{AgCl}(\text{s})$, the reactants A. gain electrons, only B. lose electrons, only C. both gain and lose electrons D. neither gain nor lose electrons 24. In the reaction

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$\text{Mg} + \text{Cl}_2 \rightarrow \text{MgCl}_2$, the correct half-reaction for the ...

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Answers Oxidation Numbers Keywords: redox, answers, oxidation, numbers Created Date: 7/16/2021 9:54:17 PM

28/6/2010 · 12. Calculate the oxidation number of chromium in each of the following. a. Cr_2O_3 b. $\text{Na}_2\text{Cr}_2\text{O}_7$ c. CrSO_4 d. chromate e. dichromate $3+$ $6+$ $2+$ $7+$ $6+$
13. Use the changes in oxidation numbers to determine which elements are oxidized and which are reduced in these reactions. (Note: it is not necessary to use balanced

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equations) a.

You are given rules for assigning oxidation numbers in Table 9.3 on page 604 of the student textbook. Plan Your Strategy Act on Your Strategy Step 1 Assign an oxidation number to each atom in the equation and determine whether it is a redox reaction. The oxidation numbers of boron and magnesium change during the reaction. It is a redox reaction.

Oxidation Number Exercise - answers Page 57 Oxidation Number Exercise Do not hand in this work sheet. When you are ready, you will be given an examination over this material. Complete the examination by yourself and hand it in to receive credit.

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Purpose: This exercise is designed to teach the student how to assign oxidation numbers.

Assigning Oxidation Numbers Answers Pg 91 PDF Assigning Oxidation Numbers Answers Pg 91 Research In Any Way. Along With Them Is This Assigning Oxidation Numbers Answers Pg 91 That Can Be Your Partner. If You Have An Internet Connection, Simply Go To BookYards And

3) Balance the following oxidation-reduction equation. $\text{FeS}_2 + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3 + \text{SO}_2$
4) Balance the following in acid. $\text{Sn}^{2+} + \text{IO}_3^- \rightarrow \text{Sn}^{4+} + \text{I}^-$
5) Balance the following in base. $\text{Bi}(\text{OH})_3 + \text{SnO}_2 \rightarrow \text{Bi} + \text{SnO}_3^{2-}$

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Answers : Part 1) 1) Oxidation - It is the loss of electrons
Reduction - It is the gain of electrons
Oxidizing agent - It is the species that ...

By referring to your periodic table, predict the oxidation number of these elements when they form ions: a. -1 b. +4 c. +4 d. +1 e. +3 f. -2 g. +2 h. +1 i. +4 j. -1 k. +2 3.

Examine each of the following reactions and decide which are 'redox' reactions.

Explain your choice: a. Cu changes to Cu^{2+} & Cl changes to Cl^- b. No changes in oxidation ...

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Redox Worksheet Answer Key Assigning

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Oxidation Numbers: Assign oxidation numbers to ...

Oxidation Number Oxidation numbers are a useful tool for determining whether a substance has been oxidized or reduced. An element that undergoes a change in oxidation number in the course of a reaction has been oxidized or reduced. Let's learn how to assign oxidation numbers.

8/5/2013 · B. reduction, only C. both oxidation and reduction D. neither oxidation nor reduction 23. In the reaction $\text{AgNO}_3(\text{aq}) + \text{NaCl}(\text{aq}) \rightarrow \text{NaNO}_3(\text{aq}) + \text{AgCl}(\text{s})$, the reactants A. gain electrons, only B. lose electrons, only C. both gain and lose electrons D. neither gain

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nor lose electrons 24. In the reaction $\text{Mg} + \text{Cl}_2 \rightarrow \text{MgCl}_2$, the correct half-reaction for the ...

G:\Obj-Oxidation & Redox Reaction - Level -II.doc Oxidation Number & Redox Reaction (Level – II) READ THE INSTRUCTIONS CAREFULLY 1. The test is of 80 minutes duration. 2. The maximum marks are 110. 3. This test consists of 35 questions. 4. For each question in Section - A, B & C you will be awarded 3 marks if you have darkened

28/6/2010 · 12. Calculate the oxidation number of chromium in each of the following. a. Cr_2O_3 b. $\text{Na}_2\text{Cr}_2\text{O}_7$ c. CrSO_4 d. chromate e. dichromate $3+ 6+ 2+ 7+ 6+$

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13. Use the changes in oxidation numbers to determine which elements are oxidized and which are reduced in these reactions. (Note: it is not necessary to use balanced equations) a.

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Oxidation-Reduction Reactions 6.2
Oxidation Numbers. 6.3 Types of Chemical
. Reactions. 6.4 Voltaic Cells. In many
important chemical reactions, electrons are
transferred from atom to atom. We are

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surrounded by these reactions, commonly called oxidation?reduction (or . redox...

Oxidation-Reduction Reactions Review Skills 6.1 An Introduction to Oxidation-Reduction Reactions Oxidation, Reduction, and the Formation of Binary Ionic Compounds Oxidation-Reduction and Molecular Compounds Special Topic 6.1: Oxidizing Agents and Aging 6.2 Oxidation Numbers Internet: Balancing Redox Reactions

UNIT 6 – REDOX REACTIONS 6 • The oxidation number of an atom is the charge that would exist on an individual atom if the bonding were completely ionic • In simple ions, the oxidation number of the atom is

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the charge on the ion: - Na^+ , K^+ , H^+ all have an oxidation number of +1 - Mg^{2+} , Ca^{2+} , Pb^{2+} all have an oxidation number of +2 - Cl^- , Br^- , I^- all have an oxidation number of -1

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