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Change • Chapter 14 277  
Section 14.1 Types of  
Mixtures pages 476 – 479  
Section Assessment 14.1  
page 479 1. Explain Use  
the properties of seawater  
to describe the

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characteristics of mixtures.  
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Answers will vary but  
might include that

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great deal of the  
information that has been  
gathered by scientists  
about matter. But,  
chemistry is ...

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This graph shows how the solubility of several solids changes with temperature.

The temperature

dependence of solubility

can be exploited to prepare supersaturated solutions of certain compounds. A

solution may be saturated



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with the compound at an  
elevated temperature  
(where the solute is more  
soluble) and subsequently  
cooled to a lower  
temperature without  
precipitating the solute.

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CHEMISTRY

HANDBOOK Matter is

made up of particles,

called atoms, that are so

small they cannot be seen

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with an ordinary light  
microscope. The struc-  
ture, composition, and  
behavior of all matter can  
be explained by atoms and  
the changes they undergo.

Solubility Versus  
Temperature This chart

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shows the solubility of various substances in water at a variety of temperatures (in degrees Celsius). Notice how NaCl's solubility is relatively constant regardless of temperature, whereas Na<sub>2</sub>SO<sub>4</sub>'s

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solubility increases  
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exponentially over 0–35  
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degrees Celsius and then  
abruptly begins to  
decrease.

2-, which pulls the  
solubility equilibrium to  
the right making  $\text{CaCO}_3$

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more soluble. In this

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supplement we will

ebooks calculate the solubility of

calcium carbonate at a

given pH taking hydrolysis

into account. This problem

is an example of a general

procedure that can be

applied to any 1:1 salt of a

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monoprotic weak acid and  
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Section 13.1 The Gas  
Laws pages 442–451  
Practice Problems page  
443 Assume that the  
temperature and the  
amount of gas are constant



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in the following problems.  
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1. The volume of a gas at  
ebooks 99.0 kPa is 300.0 mL. If

In general, solids become  
more soluble as the  
temperature increases.  
This is why sugar  
dissolves better in hot

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water than in cold water.  
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The table shows three  
examples of the solubility  
(g of ...

How Temperature  
Influences Solubility. The  
solubility of a substance is  
the amount of that

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substance that is required  
to form a saturated

solution in a given amount

of solvent at a specified

temperature. Solubility is

often measured as the

grams of solute per 100 g

of solvent. The solubility

of sodium chloride in

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water is 36.0 g per 100 g  
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water at 20°C.  
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