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Sol. (i) Here,  $q = 0$  ? ?  $U = q + w = 0 + w$  a d = w a d As no heat is absorbed by the system, the wall is adiabatic. (ii) Here,  $w = 0$ ,  $q = ?$  q As heat is taken out, the system must be having thermally conducting walls. (iii) As work is done by the system on absorbing heat, it must be a closed system.

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reaction extent 3.2.1 First Law of Thermodynamics:  $dU=dq + dw$  where U is the internal energy, q is the heat transferred to a system from the

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