

Stoichiometry Guided Study Work Answers

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Stoichiometry Calculation Practice Worksheet 1. Calculate the number of moles of NaOH that are needed to react with 500.0 g of H_2SO_4 according to the following equation: $\text{H}_2\text{SO}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$ ANS: 10.19 mol 2. Calculate the mass of NH_3 that can be produced from the reaction of 125 g of NCl_3

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according to the following equation:

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3-4 Stoichiometry Chapter 3 EXAMPLE PROBLEM: Convert Between Moles and Numbers of Atoms A sample of titanium contains 8.98×10^{25} Ti atoms. What amount of Ti, in moles, does this represent? SOLUTION: You are asked to calculate the amount (moles) of Ti in a given sample of the metal. You are given the number of atoms of Ti in the sample. Use the equality $1 \text{ mol} = 6.022 \times 10^{23}$ atoms to create ...

Stoichiometry Example C $6\text{H}_6 + \text{Br}_2 \rightarrow \text{C}_6\text{H}_5\text{Br} + \text{HBr}$
Benzene (C_6H_6) reacts with Bromine to produce bromobenzene ($\text{C}_6\text{H}_5\text{Br}$) and hydrobromic acid. If 30. g of benzene reacts with 65 g of bromine and produces 56.7 g of bromobenzene, what is the

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percent yield of the reaction? 30.g ...

Stoichiometry Calculation Practice Worksheet 1. Calculate the number of moles of NaOH that are needed to react with 500.0 g of H_2SO_4 according to the following equation: $\text{H}_2\text{SO}_4 + 2 \text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + 2 \text{H}_2\text{O}$ ANS: 10.19 mol 2. Calculate the mass of NH_3 that can be produced from the reaction of 125 g of NCl_3 according to the following equation:

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Stoichiometry grade levels; *Provide rubrics, worksheets, sentence frames, and other imaginative tools that encourage academic communication; and *Offer guiding questions to help teachers plan instruction.

Sample Study Sheet 10.3: Equation Stoichiometry Problems To get a review of the most important topics in the chapter, fill in the blanks in the Key Ideas section. Work all of the selected problems at the end of the chapter, and check your answers with the solutions provided in this chapter of the study guide. Ask for help if you need it.

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